



**HAL**  
open science

# Lipase-Catalyzed Baeyer-Villiger Oxidation of Cellulose-Derived Levoglucosenone into (S)- $\gamma$ -Hydroxymethyl- $\alpha,\beta$ -Butenolide: Optimization by Response Surface Methodology

Andreia Teixeira, Amandine Flourat, Aurelien A. M. Peru, Fanny Brunissen, Florent Allais

## ► To cite this version:

Andreia Teixeira, Amandine Flourat, Aurelien A. M. Peru, Fanny Brunissen, Florent Allais. Lipase-Catalyzed Baeyer-Villiger Oxidation of Cellulose-Derived Levoglucosenone into (S)- $\gamma$ -Hydroxymethyl- $\alpha,\beta$ -Butenolide: Optimization by Response Surface Methodology. *Frontiers in Chemistry*, 2016, 4, pp.16. 10.3389/fchem.2016.00016 . hal-01598622

**HAL Id: hal-01598622**

**<https://agroparistech.hal.science/hal-01598622>**

Submitted on 27 May 2020

**HAL** is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.



# Lipase-Catalyzed Baeyer-Villiger Oxidation of Cellulose-Derived Levoglucosenone into (S)- $\gamma$ -Hydroxymethyl- $\alpha,\beta$ -Butenolide: Optimization by Response Surface Methodology

Andreia R. S. Teixeira<sup>1,2\*</sup>, Amandine L. Flourat<sup>1,3</sup>, Aurelien A. M. Peru<sup>1,3</sup>, Fanny Brunissen<sup>1,3</sup> and Florent Allais<sup>1,4</sup>

<sup>1</sup> Chaire Agro-Biotechnologies Industrielles, AgroParisTech, Reims, France, <sup>2</sup> UMR GENIAL, AgroParisTech, Institut National de la Recherche Agronomique, Université Paris-Saclay, Massy, France, <sup>3</sup> UMR 1318 IJPB, AgroParisTech, Institut National de la Recherche Agronomique, Université Paris-Saclay, Versailles, France, <sup>4</sup> UMR 782 GMPA, AgroParisTech, Institut National de la Recherche Agronomique, Université Paris-Saclay, Thiverval-Grignon, France

## OPEN ACCESS

### Edited by:

Warwick Douglas Raverty,  
Monash University, Australia

### Reviewed by:

George Kokotos,  
University of Athens, Greece  
Victor Sans Sangorrin,  
University of Nottingham, UK

### \*Correspondence:

Andreia R. S. Teixeira  
andreia.teixeira@agroparitech.fr

### Specialty section:

This article was submitted to  
Chemical Engineering,  
a section of the journal  
Frontiers in Chemistry

Received: 23 February 2016

Accepted: 29 March 2016

Published: 19 April 2016

### Citation:

Teixeira ARS, Flourat AL, Peru AAM,  
Brunissen F and Allais F (2016)  
Lipase-Catalyzed Baeyer-Villiger  
Oxidation of Cellulose-Derived  
Levoglucosenone into  
(S)- $\gamma$ -Hydroxymethyl- $\alpha,\beta$ -Butenolide:  
Optimization by Response Surface  
Methodology. *Front. Chem.* 4:16.  
doi: 10.3389/fchem.2016.00016

Cellulose-derived levoglucosenone (**LGO**) has been efficiently converted into pure (S)- $\gamma$ -hydroxymethyl- $\alpha,\beta$ -butenolide (**HBO**), a chemical platform suited for the synthesis of drugs, flavors and antiviral agents. This process involves two-steps: a lipase-catalyzed Baeyer-Villiger oxidation of **LGO** followed by an acid hydrolysis of the reaction mixture to provide pure **HBO**. Response surface methodology (RSM), based on central composite face-centered (CCF) design, was employed to evaluate the factors effecting the enzyme-catalyzed reaction: pKa of solid buffer (7.2–9.6), **LGO** concentration (0.5–1 M) and enzyme loading (55–285 PLU.mmol<sup>-1</sup>). Enzyme loading and pKa of solid buffer were found to be important factors to the reaction efficiency (as measured by the conversion of **LGO**) while only the later had significant effects on the enzyme recyclability (as measured by the enzyme residual activity). **LGO** concentration influences both responses by its interaction with the enzyme loading and pKa of solid buffer. The optimal conditions which allow to convert at least 80% of **LGO** in 2 h at 40°C and reuse the enzyme for a subsequent cycle were found to be: solid buffer pKa = 7.5, [**LGO**] = 0.50 M and 113 PLU.mmol<sup>-1</sup> for the lipase. A good agreement between experimental and predicted values was obtained and the model validity confirmed ( $p < 0.05$ ). Alternative optimal conditions were explored using *Monte Carlo* simulations for risk analysis, being estimated the experimental region where the **LGO** conversion higher than 80% is fulfilled at a specific risk of failure.

**Keywords:** response surface methodology, reaction optimization, Bayer-Villiger bio-oxidation, lipase, enzymatic reaction, levoglucosenone

## 1. INTRODUCTION

Driving forces for the global trend of using clean renewable sources in the production of valuable chemical are the inevitable decline of fossil fuels and a more demanding legislation regarding the disposal of industrial wastes. In this context, lignocellulosic biomass is envisaged as an interesting source to produce highly valuable synthons due to its low cost and high availability. Thermochemical processing, phosphoric acid-catalyzed pyrolysis in particular, is the simplest way to efficiently convert lignocellulosic biomass into its degradation products (Huber et al., 2006; Babu, 2008). However, the use of strong acids is one of the drawbacks of such process, mainly associated to its inefficient recovery and environmental concerns related to water and air pollution (Wei et al., 2014). Furacell Process developed by Circa is claimed to be an easily scalable catalytic thermochemical process with no harmful effluents which converts renewable cellulose into a very simple mixture of easily separable products, comprising levoglucosenone (LGO), char and water reusable in the process (Court et al., 2012).

LGO is a multifunctional C<sub>6</sub>-monomer suited for organic synthesis (Miftakhov et al., 1994; Sarotti et al., 2012). Indeed, LGO is a chiral synthon that can be used in the synthesis of a wide variety of biologically active compounds such as pharmaceutical ingredients, agrochemicals, polymers, or specialty chemicals (Miftakhov et al., 1994; Budarin et al., 2011). Among them, (S)- $\gamma$ -hydroxymethyl- $\alpha,\beta$ -butenolide (aka HBO) is certainly the most interesting since it is a precursor of many drugs (Enders et al., 2002), flavors (Kawakami et al., 1990), and antiviral agents (Flores et al., 2011).

Kawakami (Kawakami et al., 1990) and Paris (Paris et al., 2013) reported the most efficient methods for HBO synthesis, being obtained high overall yields (ca. 80–90%, Figure 1).

In both methods, HBO is produced through a Baeyer-Villiger oxidation of LGO followed by an acid hydrolysis to convert formate lactone (FBO), a reaction by-product, into HBO. Kawakami et al. (Kawakami et al., 1990) used peracids (such as peracetic acid or *m*-chloroperbenzoic acid) and Me<sub>2</sub>S during 48 h followed by an HCl-mediated hydrolysis to provide pure HBO. Paris et al. (2013) developed a method to produce HBO in only 4 h using metal-based zeolites as catalysts to oxidize LGO

and Amberlyst-15 as acid resins to promote the subsequent acid hydrolysis. However, the complexity and high cost inherent to zeolites synthesis may compromise the technical and economic viability of such process (Perot and Guisnet, 1990).

The use of lipases as biocatalyst seemed to be a promising greener alternative, providing, in addition, a cost-efficient transformation. In a recent publication (Flourat et al., 2014), we reported an efficient chemo-enzymatic process for the production of HBO with high yields (> 80%). The first step involved a Baeyer-Villiger oxidation of LGO mediated by a commercial immobilized lipase from *Candida antarctica* (CAL-B, Novozyme<sup>®</sup> 435), under the presence of a solid buffer and using ethyl acetate and hydrogen peroxide as acyl donor and oxidant, respectively (Figure 2). After 2 h of reaction, the resulting mixture of FBO and HBO was hydrolyzed under acid conditions, using Amberlyst-15, to provide pure HBO.

Response surface methodology (RSM) is an effective tool for optimizing a range of processes (Montgomery, 2008) and evaluate the interactions of multiple parameters, being established in addition a prediction model based on statistics (Montgomery, 2008). RSM based on central composite face-centered (CCF) design was employed to found the conditions which allow to maximize simultaneously the conversion of LGO and the enzyme residual activity. The present work focuses on the study of reaction parameters (solid buffer pKa, LGO concentration and enzyme loading) that may affect the conversion of LGO into HBO and the enzyme recyclability (measured by the enzyme residual activity) in order to establish their relationships and, if possible, further optimize the conversion obtained in an earlier authors' publication (Flourat et al., 2014); 83%. The temperature was set at 40°C to minimize the LGO degradation (reactant specie due to its acetal moiety) and to avoid the risk of explosion linked to the *in-situ* formation of peracetic acid. The RSM model allowed explore alternative optimal conditions and using Monte Carlo simulations the risk of failure could be determined.

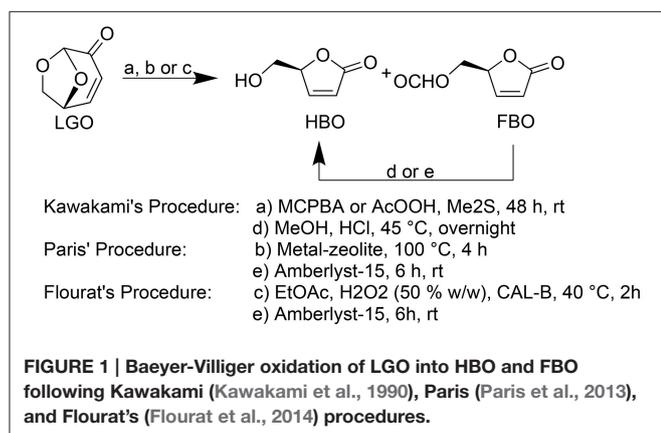
## 2. EXPERIMENTAL

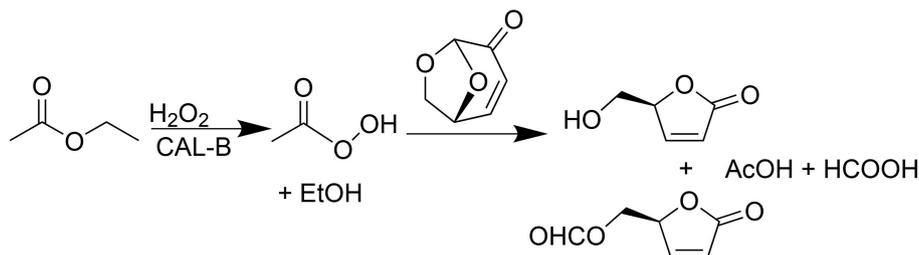
### 2.1. Reagents

Cellulose-derived levoglucosenone (LGO) was kindly provided by CIRCA Group (Knoxfield Victoria, Australia). Novozyme<sup>®</sup> 435 (CAL-B, lot no. SLBF9301 V, 9120 PLU g<sup>-1</sup>), hydrogen peroxide (50% w/w), biological solid buffers (MOPS, TAPS, CAPSO, HEPES and their sodium salt), lauric acid (99%) and 1-propanol were purchased from Sigma-Aldrich. Ethyl acetate (analytical grade), hexane (ACS reagent) and acetonitrile (HPLC grade) were purchased from Thermofisher Scientific. Ultra-pure laboratory grade water (MilliQ, 18.2 megaOhms, 25°C) was employed for HPLC analysis.

### 2.2. Baeyer-Villiger Oxidation of (-)-Levoglucosenone (LGO)

Lipase-mediated Baeyer-Villiger oxidation of LGO is schematically represented in Figure 2. Reactions were carried out in sealed erlenmeyers (50 mL) to avoid solvent evaporation. LGO (250 mg, 1 equiv.) was dissolved in different volumes of





**FIGURE 2 |** Lipase-mediated Baeyer-Villiger oxidation of LGO into HBO and FBO using AcOEt as an acyl donor and  $\text{H}_2\text{O}_2$  as an oxidant.

ethyl acetate ( $C = 0.5\text{--}1\text{ M}$ ). Biological solid buffers along with their sodium form were added ( $20\text{ mg mL}^{-1}$  each) in order to set the  $\text{pK}_a$  ( $7.5 < \text{pK}_a < 9.6$ ) and as consequence control the protonation state of the enzyme. It should be noted that unlike the organic soluble buffers, each buffer pair will set a fixed value of the relevant ionization parameter (Partridge et al., 2001). Different amounts of lipase (CAL-B) were added accordingly to those indicated in **Table 2**, followed by the addition of hydrogen peroxide (1.2 equiv., 50% w/w) at once. The mixture was incubated at  $40^\circ\text{C}$  and stirred using an orbital shaker (ThermoScientific, MaxiQ400) at 250 rpm for 2 h. Samples ( $10\ \mu\text{L}$ ) were collected and diluted with 1.5 mL of acetonitrile for LGO quantification by HPLC analysis. The conversion of LGO was defined as the molar ratio of the remaining LGO and the LGO added in the beginning.

### 2.3. Determination of Enzyme Residual Activity

After reaction, the enzyme was recovered by filtration, washed with ethyl acetate (10 mL) and hexane (10 mL), dried in an oven at  $40^\circ\text{C}$  for 1 h, and then kept in dessicator overnight and under vacuum. Three enzyme samples (ca. 14 mg each) were collected, weighed in a 20 mL vial and kept at  $60^\circ\text{C}$ . A solvent free equimolar solution of lauric acid and 1-propanol with 3% w/w of water was prepared and incubated at  $60^\circ\text{C}$ . After 1 h, 5 g of the above solution were added to the enzyme and stirred in an incubating orbital minishaker (VWR) at 400 rpm and  $60^\circ\text{C}$  for 15 min. Samples ( $2\ \mu\text{L}$ ) were collected, weighed in a GC-vial and diluted with 1 mL of hexane. Lauric acid conversion was determined by GC-MS. Unit definition: 1 PLU =  $1\ \mu\text{mol}$  of 1-propyl laurate formed per gram of enzyme per minute at  $60^\circ\text{C}$ .

### 2.4. GC-MS Method

Lauric acid (used to determine the enzyme residual activity) was quantified by a GC-MS system which consisted of an Agilent GC 5975 coupled with MS 7890 in electron impact mode with electron energy set at 70 eV and a mass range at  $m/z$  (30–350 amu). A HP5-MS capillary column (Agilent,  $30\text{ m} \times 0.25\text{ mm}$ ,  $0.25\ \mu\text{m}$ ) was used for chromatographic separation. Injection was performed at  $280^\circ\text{C}$  in split mode (40:1), being injected  $1\ \mu\text{L}$  of each sample. The oven temperature program was the following: from  $60^\circ\text{C}$  held for 1 min, then rinse until  $325^\circ\text{C}$  at  $20^\circ\text{C}/\text{min}$  with a 5 min hold. Hydrogen flow rate was set at 1.2

mL/min. The mass detector was set as follows: source and quad temperatures at 230 and  $150^\circ\text{C}$ , respectively. A calibration curve was performed each time, with pure lauric acid ( $0.2\text{--}4.5\text{ mg/mL}$ ) in hexane. Typical retention time for lauric acid and 1-propyl laurate were 7.22 min and 7.92 min, respectively.

### 2.5. HPLC Method

(-)-Levoglucosenone (LGO) was quantified by HPLC (ThermoFisher Ultimate 3000) on a Synchronis aQ column ( $250 \times 4.6\text{ mm}$ ,  $5\ \mu\text{m}$ , ThermoScientific) with Milli-Q water (solvent A) and acetonitrile (solvent B) as mobile phase. The flow rate and temperature were set at  $0.8\text{ mL}/\text{min}$  and  $30^\circ\text{C}$ , respectively. The gradient elution was as follows: isocratic at 85% A (0–5 min), from 85 to 90% A (5–10 min), isocratic at 90% A (10–15 min), from 90% A to 85% A (15–20 min). Adequate detection was obtained with a diode array detector (DAD) set at 220 nm. Injection volume in HPLC injector was set at  $10\ \mu\text{L}$ . Samples for HPLC analysis were prepared by diluting  $10\ \mu\text{L}$  of the reaction mixture in 1.5 mL of acetonitrile. The peaks of LGO were identified and quantified using a standard curve prepared in acetonitrile. Typical retention time for HBO, FBO and LGO were 3.73, 3.87, and 8.40 min, respectively.

### 2.6. Experimental Design and Statistical Analysis

RSM, based on a 3-factor-3-level CCF design, was employed to determine the parameters affecting the Baeyer-Villiger oxidation of LGO as well as to found the optimal set of conditions. CCF is a good choice from a practical point of view, even inducing some correlations between the quadratic terms. These correlations will induce a slight increase in the confidence intervals but the model will still be able to estimate the quadratic effects.

A short reaction time of 2 h was set in order to minimize the residence time of the enzyme and limit the contact with the inhibitors as well as to minimize the residence time of LGO inside the reactor and, thus, minimizing LGO degradation. The temperature was set at  $40^\circ\text{C}$  for the same reason but also to minimize the risk of explosion linked to the presence of peracetic acid (intermediary specie).

**Table 1** presents the independent variables ( $x_i$ ), levels and experimental design in terms of uncoded and coded (transformation of each studied real value into coordinates inside

**TABLE 1 | Independent variables and levels used for CCF design.**

Variables	Level		
	-1	0	1
Solid buffer pka	MOPS (7.2)	TAPS (8.4)	CAPSO (9.6)
Enzyme loading (PLU.mmol <sup>-1</sup> )	55	170	285
<b>LGO</b> concentration (M)	0.5	0.75	1

a scale with dimensionless values). The variation ranges of each variable were fixed taking into account specific constraints:

- Following the manufacture (Novozymes, 2015) the optimal pka for the lipase CAL B lies between 5 and 9. However, the pka buffer cannot be acid (< 7) in order to avoid the degradation of **LGO** (acetal moieties reactive under such conditions). Therefore, the effect of pka in the enzyme protonation state was assessed with MOPS (pka = 7.2) and CAPSO (pka = 9.6), using TAPS (pka = 8.4) as a central point.
- A lower solvent volume to obtain higher **LGO** concentrations implies the presence of inhibitors (peracetic acid, acetic acid and formic acid) at higher concentrations. Moreover, the use of high **LGO** concentrations (> 1 M) leads to a decrease in the **LGO** conversion and (mainly) in the enzyme activity (Flourat et al., 2014). Instead, a low concentration of substrates implies the use of large quantities of solvents, which may compromise the economic sustainability of a process. Thus, **LGO** concentration was varied between 0.5 and 1 M.
- Enzymes must be used at catalytic quantities to assure the economic sustainability of the process. Thus, the enzyme loading was varied between 55 and 285 PLU.mmol<sup>-1</sup>, which corresponds to 2 and 10 % (w/w), respectively.

**Table 2** presents the runs set by a CCF design and the respective experimental responses obtained for **LGO** conversion ( $Y_1$ ) and enzyme residual activity ( $Y_2$ ). CCF design consists of 8 factorial points, 6 axial points (two axial points on the axis of each design variable at a distance 1 from the design point) and 3 central points (Eriksson et al., 2008), making a total of 17 runs. Duplicates in each point were performed in order to obtain a more precise model (total of 34 runs). To avoid bias, all runs were performed in a totally random order. Additional runs (intermediary points) were performed to validate the model and later added to the experimental data for model refining.

In order to find a suitable approximation for the true functional relationship between independent variables and the response surface, a second-order polynomial Equation (1) was used, being expressed as:

$$Y = \beta_0 + \sum_{i=1}^3 \beta_{ki}x_i + \sum_{i=1}^3 \beta_{kii}x_i^2 + \sum_{i=1}^2 \sum_{j=i+1}^3 \beta_{kij}x_i x_j \quad (1)$$

where  $Y_i$  represents the response  $i$  (**LGO** conversion and enzyme residual activity in this study),  $x_i$  are the coded independent variables,  $\beta_0$  is a constant coefficient, and  $\beta_{ki}$ ,  $\beta_{kii}$ , and  $\beta_{kij}$  are the linear, quadratic, and interaction coefficients, respectively.

The variable levels  $X_i$  were scaled and centered (coded) according to the equation below such that  $X_0$  corresponded to the central value:

$$x_i = \frac{X_i - X_0}{\Delta X_i} \quad i = 1, 2, 3, \dots, k \quad (2)$$

where  $x_i$  is the dimensionless value of an independent variable,  $X_i$  is the real value of an independent variable,  $X_0$  is the real value of an independent variable at the center point, and  $\Delta X_i$  is the step change.

Modde v.10.1 software (Umetrics AB, Sweden) was used to generate the CCF design and analyze experimental data by RSM. Regression coefficients were determined by multiple linear regression (MLR). The significant parameters in the model were found by analysis of their  $p$ -value. The model validation was based on the variance (ANOVA) for each response, namely, by the analysis of  $R^2$ ,  $Q^2$ , and *lack of fit* (LOF) test.  $R^2$  measures how well the regression model fits the experimental data,  $Q^2$  shows an estimate of the future prediction precision, and, LOF assess whether the models error is comparable to the replicate error.

### 3. RESULTS AND DISCUSSION

#### 3.1. Preliminary Study

In our previous study (Flourat et al., 2014) we reported the ability of commercial immobilized *Candida antarctica* lipase (Novozyme® 435) to oxidize **LGO** (**Figure 2**). High conversions of **LGO** (> 83%) were obtained after 2 h, using a low enzyme loading (113 PLU.mmol<sup>-1</sup>) and a **LGO** concentration of 0.75 M. Such mild conditions allowed to reuse the enzyme for a second reaction cycle without a significant loss of enzyme activity.

The optimization was conducted using the One Variable At a Time (OVAT) method, being optimized the temperature, **LGO** concentration, enzyme loading, nature of buffer (solid/liquid and pka) as well as the nature and quantity of the oxidizing agent.

Results showed that at low temperatures (< 40°C) the final **LGO** conversion is compromised, being, although, similar at 40 and 60°C (after 8 h of reaction). However, there are other reasons to remain at 40°C than an obvious energy saving: the **LGO** degradation (due to its acetal moiety) as well as the risk of explosion linked to the presence of peracetic acid (intermediary specie). For all this reasons, the temperature was set at 40°C in this work even having to ignore interactions with other variables.

Moreover, it was showed that a low-water media (resulted in using solid buffers) led to higher conversion of **LGO** probably due to a higher stability of the enzyme and inhibition of secondary reactions. Unfortunately in this reaction, the water activity cannot be actively controlled due to high water content of commercial hydrogen peroxide solution (50% w/w).

As summary, the previous study (Flourat et al., 2014) suggested that only three variables had significant effect on the **LGO** conversion. These variables were the buffer pka ( $x_1$ ), the enzyme loading ( $x_2$ ) and the **LGO** concentration ( $x_3$ ). The level values of variables were chosen in such a way that their limits were as wide as possible and that contains the optimal values determined using the OVAT method (**Table 1**). The levels of the three other variables with a small effect on the oxidation reaction

**TABLE 2 | Central composite face-centered (CCF) design and experimental responses.**

Run/Replicate <sup>a</sup>	Solid buffer (pKa)	Enzyme loading (PLU.mmol <sup>-1</sup> of LGO)	[LGO] (M)	LGO conversion (%)	Enzyme residual activity (%)
	x <sub>1</sub>	x <sub>2</sub>	x <sub>3</sub>	Y <sub>1</sub>	Y <sub>2</sub>
1/18	MOPS (7.2)	55	0.5	57.4/68.0	88.5/82.0
2/19	CAPSO (9.6)	55	0.5	79.9/79.7	63.6 <sup>b</sup> /37.6
3/20	MOPS (7.2)	285	0.5	81.3/87.0	85.3/81.0
4/21	CAPSO (9.6)	285	0.5	87.5/90.8	52.7/57.3
5/22	MOPS (7.2)	55	1.0	74.8/74.5	70.3/76.9
6/23	CAPSO (9.6)	55	1.0	47.3/63.9	60.2/60.7
7/24	MOPS (7.2)	285	1.0	81.5/91.6	51.2/44.8
8/25	CAPSO (9.6)	285	1.0	87.2/90.4	69.2/52.4 <sup>b</sup>
9/26	MOPS (7.2)	170	0.75	84.4/87.2	72.4/73.4
10/27	CAPSO (9.6)	170	0.75	90.0/88.9	68.9/67.0
11/28	TAPS (8.4)	55	0.75	73.9/82.0	73.1/63.1
12/29	TAPS (8.4)	285	0.75	67.8 <sup>b</sup> /91.3	74.8/50.6
13/30	TAPS (8.4)	170	0.5	86.6/87.6	61.8/61.1
14/31	TAPS (8.4)	170	1.0	88.4/89.5	12.6 <sup>b</sup> /67.5
15/32	TAPS (8.4)	170	0.75	87.9/90.2	65.7/68.5
16/33	TAPS (8.4)	170	0.75	93.8/90.3	74.9/74.1
17/34	TAPS (8.4)	170	0.75	92.4/89.9	76.8/75.8
<b>ADDITIONAL RUNS</b>					
35/36	MOPS (7.2)	152	0.70	78.0/80.8	78.2/79.4
37/38	CAPSO (9.6)	80	0.94	72.5/73.6	70.0/71.6
39	HEPES (7.5)	120	0.65	79.7	76.5

<sup>a</sup>Runs performed in a totally random order.

<sup>b</sup>Outlier: observation point that is distant from other observations. Excluded from the data set.

were fixed as follow: temperature: 40°C, reaction time: 2 h, nature and quantity of the oxidizing agent: hydrogen peroxide (H<sub>2</sub>O<sub>2</sub>) at 1.2 equiv. (relative to LGO).

### 3.2. Assessing the Design Orthogonality

The condition number is a parameter that assesses the sphericity of the design, thus, the orthogonality. Formally, the condition number is the ratio of the largest and the smallest singular values of the X-matrix, that is, the matrix of the factors extended with higher order terms. As a thumb rule for an optimization (Eriksson et al., 2008), the condition number should be lower than 8. The condition number obtained in our model was 4.1 for LGO conversion and 1.6 for the enzyme activity. Such low values demonstrate well the orthogonality of the selected design, thus, the adequacy of the design selected.

As can be seen in **Table 2**, additional runs were performed. Such procedure allowed improving the model quality at an exploratory level, however, without compromising the model orthogonality.

### 3.3. Optimization by Design of Experiments (DOE)

The analysis of experimental data through DOE consists of four primary stages. The first stage, *evaluation of raw data*, focuses on

identifying regularities and peculiarities in the experimental data. The second stage, *regression analysis*, involves the calculation of the model linking the variables and response(s) together, and is followed by a third stage, the *model interpretation*. Finally, in the fourth stage, *use of regression model*, the model obtained is used to predict the optimal experimental conditions to maximize/minimize the response(s).

#### 3.3.1. Testing the Normality for Raw Data Evaluation

In regression analysis, it is advantageous if data of a response are normally distributed. This improves the efficiency of data analysis, and enhances model validity and inferential reliability. It is not recommended to apply regression analysis to a response with heavy tails as originally observed for LGO conversion (Y<sub>1</sub>) and enzyme residual activity (Y<sub>2</sub>) responses (Figure S1), since that would correspond to assigning the extreme measurement an undue influence in the modeling (Eriksson et al., 2008). Normality can be assessed to some extent by obtaining skewness and kurtosis values. Skewness is a measure of the asymmetry of the probability distribution of a random variable about its mean. Kurtosis is a measure of the “peakedness” of the probability distribution of a random variable. In other words, kurtosis indicates how tall and sharp the central peak is, relative to that of a standard bell curve. As a general rule of thumb the skewness

and kurtosis values should range between  $-0.5$  and  $0.5$  in a normal distribution (Montgomery, 2008). If these values are not included within such range, a transformation of the response should be performed (see Supplementary Material for more details). **Table 3** compares these values for both responses before and after negative logarithmic transformation.

As shown in **Table 3**, skewness and kurtosis values are within the range after response transformation. In this transformation, each measured value is subtracted from the maximum value (100, for variables expressed in percentages), and then the negative logarithm is applied (see Supplementary Material for more details, including other tools for measure the raw data normality).

### 3.3.2. Regression Analysis

The next stage consists of fitting the second-order polynomial Equation (1) to the experimental data (**Table 2**) and determining the significant coefficients for each response.

Designs with a low condition number mean having low correlations among the terms in the model. As can be seen in **Table 4**, significant correlations are only observed between variables and responses (values highlighted in bold). An exception for the correlations between quadratic terms, as a direct consequence of the design selection (CCF). However, it should be highlighted that these correlations induce a slight increase in the confidence intervals but the model will still be able to estimate the quadratic effects.

**TABLE 3 | Skewness and kurtosis values for LGO conversion ( $Y_1$ ) and enzyme residual activity ( $Y_2$ ) responses before and after a negative logarithmic transformation.**

Test	Before		After	
	$Y_1$	$Y_2$	$Y_1$	$Y_2$
Skewness	-1.2	-0.67	-0.46	0.32
Kurtosis	1.1	0.32	-0.43	0.50

**TABLE 4 | Correlation matrix.**

	$x_1$	$x_2$	$x_3$	$x_1 \times x_1$	$x_2 \times x_2$	$x_3 \times x_3$	$x_1 \times x_2$	$x_1 \times x_3$	$x_2 \times x_3$	$Y_1$	$Y_2$
$x_1$	1	-0.048	0.104	0.004	0.097	0.100	-0.086	0.048	-0.072	0.088	<b>-0.505</b>
$x_2$	-0.048	1	-0.070	0.010	0.040	0.015	0.051	-0.067	0.050	<b>0.674</b>	-0.182
$x_3$	0.104	-0.070	1	0.010	0.026	0.040	-0.072	0.051	-0.075	-0.052	-0.210
$x_1 \times x_1$	0.004	0.010	0.010	1	<b>0.492</b>	<b>0.382</b>	-0.041	0.060	-0.043	<b>-0.463</b>	0.078
$x_2 \times x_2$	0.097	0.040	0.026	<b>0.492</b>	1	<b>0.578</b>	-0.037	-0.035	-0.035	<b>-0.423</b>	-0.179
$x_3 \times x_3$	0.100	0.015	0.040	<b>0.382</b>	<b>0.578</b>	1	0.011	-0.011	-0.103	<b>-0.361</b>	-0.120
$x_1 \times x_2$	-0.086	0.051	-0.072	-0.041	-0.037	0.011	1	-0.084	0.047	0.123	0.274
$x_1 \times x_3$	0.048	-0.067	0.051	0.060	-0.035	-0.011	-0.084	1	-0.020	<b>-0.303</b>	<b>0.544</b>
$x_2 \times x_3$	-0.072	0.050	-0.075	-0.043	-0.035	-0.103	0.047	-0.020	1	0.149	-0.067
$Y_1$	0.088	<b>0.674</b>	-0.052	<b>-0.463</b>	<b>-0.423</b>	<b>-0.361</b>	0.123	<b>-0.303</b>	0.149	1	-0.240
$Y_2$	<b>-0.505</b>	-0.182	-0.210	0.078	-0.179	-0.120	0.274	<b>0.544</b>	-0.067	-0.240	1

Significant correlations are identified in bold.

The standard approach for selecting significant coefficients for each response is based on their  $p$ -value (Eriksson et al., 2008; Montgomery, 2008): a  $p$ -value lower than 0.05 means that the coefficient is significant (**Tables 5, 6**, values highlighted in bold). The coefficients centered and scaled (Coeff. SC, referring to the coded  $-1$  to  $+1$  unit) as well as their standard error (Std. Err.),  $p$ -value and confidence interval at 95% (CI) are listed in the **Tables 5, 6** for the LGO conversion and the enzyme residual activity models, respectively.

For the model of LGO, following the  $p$ -value analysis of each coefficient, LGO concentration ( $x_3$ ), the respective quadratic term ( $x_3 \times x_3$ ), the interaction between pka, the enzyme loading ( $x_1 \times x_2$ ) as well as the interaction between enzyme loading and LGO concentration ( $x_2 \times x_3$ ) are not significant terms. For the model of enzyme residual activity, the enzyme loading ( $x_2$ ), the LGO concentration ( $x_3$ ), all quadratic terms as well as the interaction between enzyme loading and LGO concentration ( $x_2 \times x_3$ ) are not significant terms. As result, the fitted model for each response, expressed in coded variables (scaled and centered), may be represented by the following equations (see in Supplementary Material the model equations in uncoded units, as well as their 3D graphical representation, Figure S6):

$$-\log_{10}(100 - Y_1) = -0.993 + 0.052x_1 + 0.208x_2 - 0.123x_1^2 - 0.113x_2^2 - 0.090x_1x_3 \quad (3)$$

$$-\log_{10}(100 - Y_2) = -1.50 - 0.111x_1 + 0.060x_1x_2 + 0.131x_1x_3 \quad (4)$$

Analysis of variance (ANOVA, **Tables 7, 8**) indicates that second-order polynomial model is adequate to represent the actual relationships between the responses (LGO conversion,  $Y_1$ , and enzyme residual activity,  $Y_2$ ) and the significant variables ( $p < 0.05$ ). Satisfactory coefficient of determination ( $R^2 > 0.5$ , using *least squares* method) and coefficient of cross-validation ( $Q^2 > 0.5$ , using the leave-one-out cross-validation) were obtained, showing the *goodness of fit* and the *goodness of prediction*, respectively. Moreover,  $F$ -tests performed in ANOVA assessing

**TABLE 5 | Model coefficients [centered and scaled (SC)], their standard error (Std. Err.), *p*-value and confidence interval at 95% (CI) for the LGO conversion model.**

Yield	Coeff. SC	Std. Err.	<i>p</i> -value	CI 95% (±)
Constant	0.992	0.029	<b>3.62</b> × 10 <sup>-24</sup>	0.061
<i>x</i> <sub>1</sub>	0.050	0.020	<b>0.020</b>	0.041
<i>x</i> <sub>2</sub>	0.201	0.022	<b>6.08</b> × 10 <sup>-10</sup>	0.045
<i>x</i> <sub>3</sub>	0.004	0.021	0.841	0.043
<i>x</i> <sub>1</sub> * <i>x</i> <sub>1</sub>	-0.125	0.039	<b>3.59</b> × 10 <sup>-3</sup>	0.081
<i>x</i> <sub>2</sub> * <i>x</i> <sub>2</sub>	-0.119	0.044	<b>0.011</b>	0.089
<i>x</i> <sub>3</sub> * <i>x</i> <sub>3</sub>	-0.057	0.039	0.165	0.081
<i>x</i> <sub>1</sub> * <i>x</i> <sub>2</sub>	0.020	0.024	0.417	0.049
<i>x</i> <sub>1</sub> * <i>x</i> <sub>3</sub>	-0.082	0.024	<b>1.74</b> × 10 <sup>-3</sup>	0.048
<i>x</i> <sub>2</sub> * <i>x</i> <sub>3</sub>	0.028	0.024	0.245	0.049

Significant *p*-values (< 0.05) are identified in bold.

**TABLE 6 | Model coefficients [centered and scaled (SC)], their standard error (Std. Err.), *p*-value and confidence interval at 95% (CI) for the enzyme residual activity model.**

Yield	Coeff. SC	Std. Err.	<i>p</i> -value	CI 95% (±)
Constant	-1.48	0.029	<b>2.49</b> × 10 <sup>-28</sup>	0.060
<i>x</i> <sub>1</sub>	-0.090	0.020	<b>1.73</b> × 10 <sup>-4</sup>	0.043
<i>x</i> <sub>2</sub>	-0.032	0.022	0.158	0.046
<i>x</i> <sub>3</sub>	-0.039	0.022	0.090	0.046
<i>x</i> <sub>1</sub> * <i>x</i> <sub>1</sub>	0.035	0.039	0.377	0.080
<i>x</i> <sub>2</sub> * <i>x</i> <sub>2</sub>	-0.011	0.043	0.790	0.088
<i>x</i> <sub>3</sub> * <i>x</i> <sub>3</sub>	-0.040	0.042	0.343	0.086
<i>x</i> <sub>1</sub> * <i>x</i> <sub>2</sub>	0.059	0.025	<b>0.025</b>	0.051
<i>x</i> <sub>1</sub> * <i>x</i> <sub>3</sub>	0.123	0.024	<b>2.67</b> × 10 <sup>-5</sup>	0.050
<i>x</i> <sub>2</sub> * <i>x</i> <sub>3</sub>	-0.027	0.025	0.284	0.051

Significant *p*-values (< 0.05) are identified in bold.

the significance of the regression model (*p* < 0.05) and the lack of fit (*p* > 0.05) showed statistical significance of both models and a similar magnitude of replicate errors (no lack of fit).

### 3.3.3. Model Interpretation

Model interpretation plays an important role in DOE. Model coefficients given by the Equations (3) and (4) are unscaled in order to enable the direct application of equations without calculating the corresponding code level of each variable. However, when scaled and centered (referring to the coded -1 to +1 unit), the coefficients values are useful for model interpretation (see the values in the Table 5, 6).

Figure 3 shows the graphical representation of the coefficient scaled and centered of the models of LGO conversion and enzyme residual activity, where positive values indicate a synergistic effect on the response while the negative indicate an antagonistic effect.

As can be observed, enzyme loading (*x*<sub>2</sub>) is the most significant variable affecting positively the conversion of LGO (*Y*<sub>1</sub>), being observed as well a negative quadratic effect mainly

by the solid buffer *p**K*<sub>a</sub> (*x*<sub>1</sub>) and enzyme loading (*x*<sub>2</sub>). LGO concentration (*x*<sub>3</sub>) affects the conversion of LGO only due to its interaction with the *p**K*<sub>a</sub> variable (*x*<sub>1</sub>*x*<sub>3</sub>). These results were expected, since increasing the enzyme loading more enzyme units are available to convert the substrate, thus, higher is the conversion of LGO. Of course, there are economic restraints in using high enzyme loading which should be taken into account in the optimization. On the other side, the effect of *p**K*<sub>a</sub> (equivalent to the pH in aqueous solutions) on the enzyme activity is widely reported in literature.

The model of the enzyme residual activity (*Y*<sub>2</sub>) is more complex. The solid buffer *p**K*<sub>a</sub> (*x*<sub>1</sub>) proved to have a significant negative effect as well as interacts with the enzyme loading and the LGO concentration (*x*<sub>1</sub>*x*<sub>2</sub> and *x*<sub>1</sub>*x*<sub>3</sub>). The complexity of the enzyme residual activity and existence of interactions is understandable and expected when taking into account all possible mechanisms of activation and inactivation of an enzyme.

The relationships between variables and responses can be better understood by examining the contour plots (Figure 4) generated from the models predicting the LGO conversion and enzyme residual activity responses (Equations 3 and 4, respectively). The contour plots can provide the contour lines of the independent variables (*x*<sub>*i*</sub>) that have the same response value *Y*<sub>*i*</sub>.

Contour plot of LGO conversion (Figure 4A) shows that high values (> 90%) can be reached at higher enzyme loading (175 PLU.mmol<sup>-1</sup>) and *p**K*<sub>a</sub>'s (>8.4). However, these conditions corresponds to the lowest enzyme residual activity (Figure 4B). The maximal value for such response (> 85%) was observed at low *p**K*<sub>a</sub> (MOPS, 7.2), low enzyme loading (< 100 PLU.mmol<sup>-1</sup>). Therefore, a trade-off between responses should be placed in order to attain the optimal conditions.

### 3.3.4. Use of Regression Model for Attaining Optimal Conditions

The optimal conditions were determined using a Nelder-Mead Simplex algorithm. This method computes the variable values (*x*<sub>1</sub>, *x*<sub>2</sub>, and *x*<sub>3</sub>) that minimizes simultaneously the normalized distance to target values of responses and DPMO (Defects Per Million Opportunities outside specifications). DPMO gives information about robustness to small disturbances introduced by the precision specified for the factors. The optimal conditions to obtain at least 80% of LGO conversion (*Y*<sub>1</sub>) and enzyme residual activity (*Y*<sub>2</sub>) were found to be: *p**K*<sub>a</sub> = 7.5 (HEPES), LGO concentration = 0.50 M and enzyme loading = 194 PLU.mmol<sup>-1</sup>. The advantage of using HEPES as solid buffer, in contrast to MOPS (*p**K*<sub>a</sub> = 7.2), is its easier recovery for reuse in a consecutive batch, since it does not form a gel.

Figure 5 shows the Sweet Spot plot when setting 80% as the minimum value for both responses. As can be observed, the criteria (area highlighted in green) were exclusively met at low *p**K*<sub>a</sub> (HEPES, *p**K*<sub>a</sub> = 7.5) and LGO concentration (< 0.57 M).

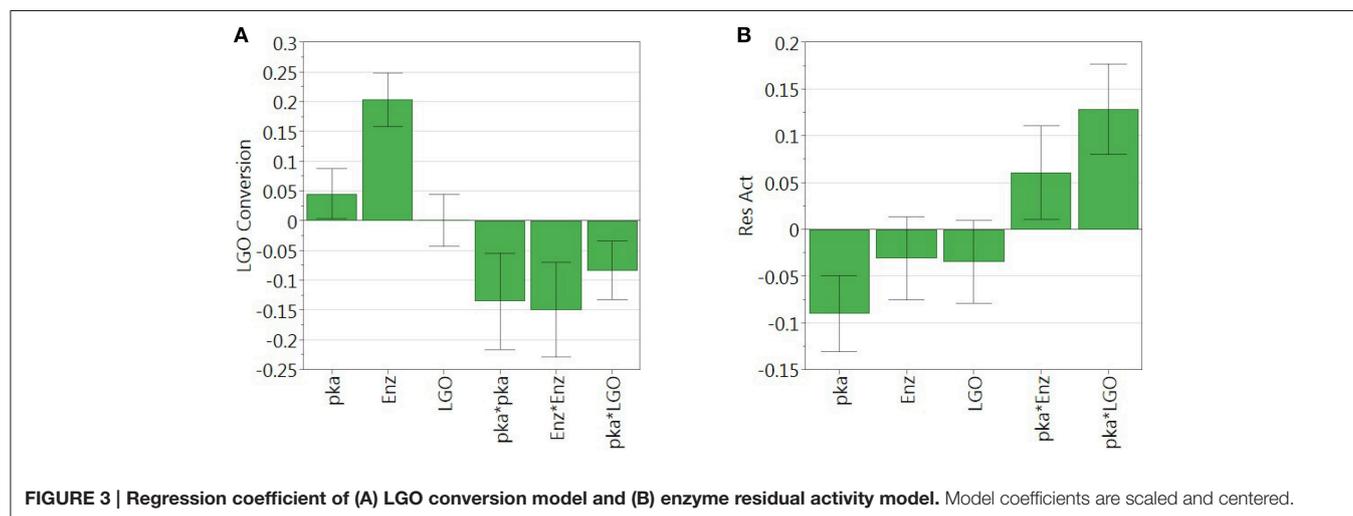
Alternative optimal conditions can be explored using a design space (DS) plot. This plot uses Monte Carlo simulations for risk analysis, estimating the volume in the experimental design region where it can be expected that all specifications are fulfilled at a specific risk level. Figure 6 shows the DS plot using as

**TABLE 7 | Analysis of variance (ANOVA) for the quadratic polynomial model fitted to LGO conversion,  $Y_1$ .**

Source of variation	Degrees of freedom	Sum of squares	Mean square	Standard deviation	Significance
Regression	6	1.498	0.249	0.500	0.000 <sup>a</sup>
Residuals	31	0.312	0.010	0.100	
Lack of fit (model error)	11	0.133	0.012	0.110	0.270 <sup>b</sup>
Pure error (replicate error)	20	0.179	0.008	0.095	
$R^2 / R^2_{adj.}$ <sup>c</sup>	0.828/0.794				
$Q^2$	0.733				

<sup>a</sup>Significant at the level 95%.<sup>b</sup>No lack of fit.<sup>c</sup> $R^2$  adjusted for degree of freedom.**TABLE 8 | Analysis of variance (ANOVA) for the quadratic polynomial model fitted to enzyme residual activity,  $Y_2$ .**

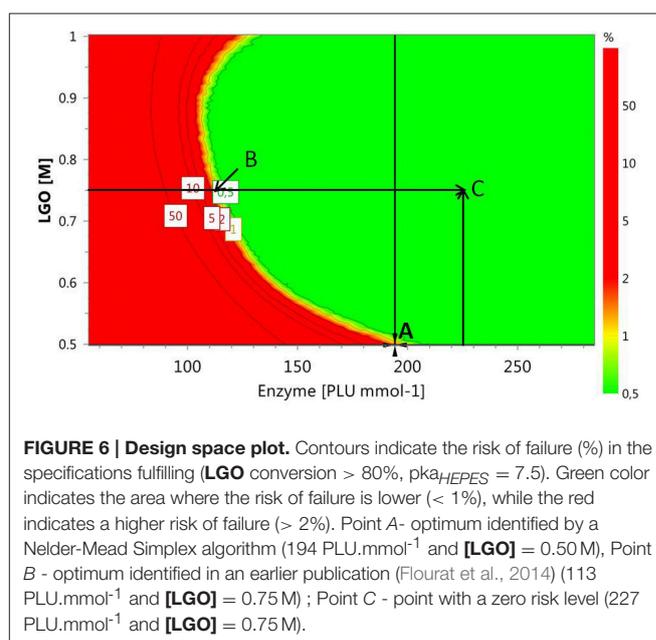
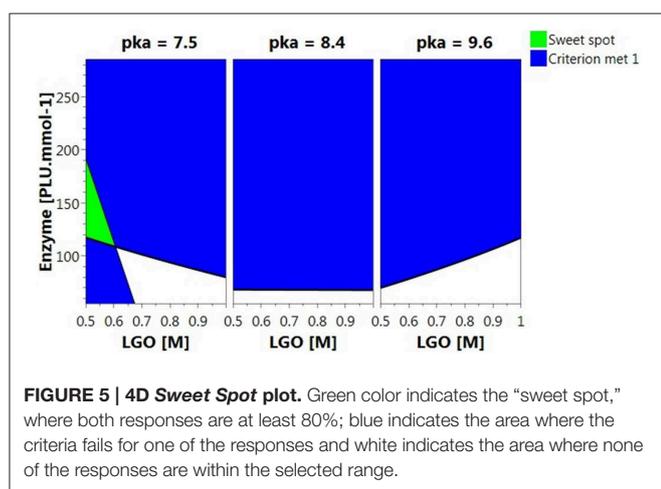
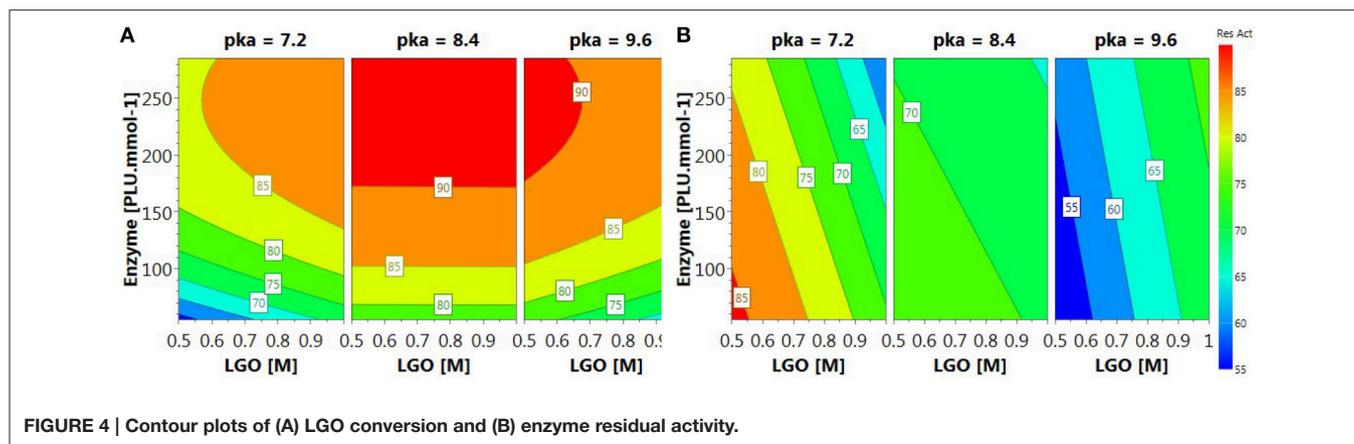
Source of variation	Degrees of Freedom	Sum of Squares	Mean square	Standard Deviation	Significance
Regression	5	0.583	0.117	0.342	0.000 <sup>a</sup>
Residuals	31	0.282	0.009	0.095	
Lack of fit (model error)	12	0.144	0.012	0.109	0.162 <sup>b</sup>
Pure error (replicate error)	19	0.139	0.007	0.085	
$R^2 / R^2_{adj.}$ <sup>c</sup>	0.674 / 0.621				
$Q^2$	0.50				

<sup>a</sup>Significant at the level 95%.<sup>b</sup>No lack of fit.<sup>c</sup> $R^2$  adjusted for degree of freedom.

specification a high **LGO** conversion ( $Y_1 > 80\%$ ). However, it should be taken into account that the lower enzyme residual activity, the greater **LGO** conversion ( $>80\%$ ). Three points were identified using this plot. Point A is the optimum identified by a Nelder-Mead Simplex algorithm, Point B uses a lower enzyme loading to convert at least 80% of **LGO** (both with a risk level

between 1 and 2%) and Point C illustrates a condition where no risk is taken to obtain 80% as minimum value for both responses, in spite of using a high enzyme loading.

The advantage of using the conditions expressed by the point B is the low enzyme loading (113 PLU.mmol<sup>-1</sup>) necessary to obtain an acceptable **LGO** conversion (80%), which represents



an economic saving and increases the viability of the process at industrial scale. The point B coincides with the optimal point identified by the authors in an earlier publication (Flourat et al., 2014), however, using the OVAT method.

### 3.4. External Model Validation

The internal validity of the predicting model was assessed by the  $Q^2$  coefficient, obtained by leave-one-out cross-validation. Points A, B, and C were used for an additional external validation.

As shown in Table 9, the experimental responses are in agreement to the ones predicted by the model, thus, the model validity can be inferred within the region (ranges of variables) studied in this work.

### 3.5. Enzyme Recyclability and HBO Production

The possibility of reusing the enzyme under point B conditions at multigram scale was shown by the authors in an early publication (Flourat et al., 2014), being obtained high conversion of LGO in the first two cycles (> 80% per cycle). Scale up of this process was performed with 10 g of LGO per oxidation cycle. Under such conditions, conversions of LGO were comparable to those

observed at the 500 mg scale (1st cycle: 83 and 81% respectively). The reaction mixture was then combined and subjected to the acid hydrolysis. Finally, after concentration to dryness, the crude mixture was easily purified by column chromatography and provided pure HBO in 67% yield (overall yield for the two 2-h oxidation cycles). <sup>1</sup>H and <sup>13</sup>C NMR spectra can be found in our early publication (Flourat et al., 2014).

## 4. CONCLUSION

RSM has proven to be adequate for the optimization of the enzymatic Baeyer-Villiger oxidation of LGO, providing, in addition, a better understanding of the individual and mutual effects of buffer pKa, enzyme loading and LGO concentration on the overall reaction efficiency (measured as conversion of LGO) as well as on the enzyme recyclability (measured as enzyme residual activity). Enzyme loading and solid buffer pKa were

TABLE 9 | Runs performed for external model validation.

Run	Factor			Estimated response		Experimental response	
	$x_1$ (pKa)	$x_2$ (PLU.mmol <sup>-1</sup> )	$x_3$ (M)	$Y_1$	$Y_2$	$Y_1$	$Y_2$
A	HEPES (7.5)	194	0.50	83.6 ± 4.1	80.6 ± 3.4	79.5	80.4
B	HEPES (7.5)	113	0.75	82.9 ± 2.5	76.1 ± 2.8	84.0	74.3
C	HEPES (7.5)	227	0.75	89.5 ± 4.5	71.6 ± 3.5	94.0	69.4

found to be important variables to the LGO conversion while for the enzyme residual activity, only the pKa and all their interactions were significant. LGO concentration influences both responses by their interaction with the enzyme loading and solid buffer pKa.

An antagonist effect of the variables on both responses was observed, thus, being necessary to establish a compromise to attain the optimal conditions. Such conditions were found to be: solid buffer pKa = 7.5, [LGO] = 0.75 M and 113 PLU.mmol<sup>-1</sup> for the lipase. Under these conditions, a high conversion (> 80%) was obtained in two consecutive batch at multigram scale.

The statistical models obtained by RSM for each response, represented by Equations (3) and (4), enable the prediction of LGO conversion and enzyme residual activity, respectively, at different conditions of pKa, enzyme loading and LGO concentration. The validity of the model was confirmed ( $p < 0.05$ ), being observed a good agreement between experimental and predicted values.

## AUTHOR CONTRIBUTIONS

FA conceived and supervised the research project. AT performed the experimental design and its statistical analysis. AF and AP

carried-out all the enzymatic reaction and quantified LGO by HPLC. AT done all the enzymatic residual activity tests, while FB quantified by GC-MS. AT drafted the manuscript, all authors were involved in revising it; FA supervised its preparation. All authors have approved and are accountable for the final version of the manuscript.

## FUNDING

The authors are grateful to Région Champagne-Ardenne, Conseil Général de la Marne and Reims Métropole for financial support.

## ACKNOWLEDGMENTS

The authors are grateful to Circa Group for providing industrial grade levoglucosenone.

## SUPPLEMENTARY MATERIAL

The Supplementary Material for this article can be found online at: <http://journal.frontiersin.org/article/10.3389/fchem.2016.00016>

## REFERENCES

- Babu, B. V. (2008). Biomass pyrolysis: a state-of-the-art review. *Biofuels Bioprod. Bioref.* 2, 393–414. doi: 10.1002/bbb.92
- Budarin, V. L., Shuttleworth, P. S., Dodson, J. R., Hunt, A. J., Lanigan, B., Marriott, R., et al. (2011). Use of green chemical technologies in an integrated biorefinery. *Energy Environ. Sci.* 4, 471–479. doi: 10.1039/C0EE00184H
- Court, G., Lawrence, C., Raverty, W., and Duncan, A. (2012). *Method for Converting Lignocellulosic Materials into Useful Chemicals*. EP Patent Application EP20100793418.
- Enders, D., Lausberg, V., Del Signore, G., and Berner, O. M. (2002). A general approach to the asymmetric synthesis of lignans: (–)-Methyl piperitol, (–)-sesamin, (–)-aschantin, (+)-yatein, (+)-dihydroclusin, (+)-burseran, and (–)-isostegane. *Synthesis* 33, 515–522. doi: 10.1055/s-2002-20967
- Eriksson, L., Johansson, E., Kettaneh-Wold, N., Wikström, C., and Wold, S. (2008). *Design of Experiments: Principles and Applications, 3rd Edn*. Umeå: Umetrics Academy.
- Flores, R., Rustullet, A., Alibés, R., Álvarez Larena, A., de March, P., Figueredo, M., et al. (2011). Synthesis of purine nucleosides built on a 3-oxabicyclo[3.2.0]heptane scaffold. *J. Org. Chem.* 76, 5369–5383. doi: 10.1021/jo200775x
- Flourat, A. L., Peru, A. A. M., Teixeira, A. R. S., Brunissen, F., and Allais, F. (2014). Chemo-enzymatic synthesis of key intermediates (s)- $\gamma$ -hydroxymethyl- $\alpha$ ,  $\beta$ -butenolide and (s)- $\gamma$ -hydroxymethyl- $\gamma$ -butyrolactone via lipase-mediated Baeyer-Villiger oxidation of levoglucosenone. *Green Chem.* 17, 404–412. doi: 10.1039/C4GC01231C
- Huber, G. W., Iborra, S., and Corma, A. (2006). Synthesis of transportation fuels from biomass: chemistry, catalysts, and engineering. *Chem. Rev.* 106, 4044–4098. doi: 10.1021/cr068360d
- Kawakami, H., Ebata, T., Koseki, K., Matsumoto, K., Matsushita, H., Naoi, Y., et al. (1990). Stereoselectivities in the coupling reaction between silylated pyrimidine bases and 1-halo-2,3-dideoxyribose. *Heterocycles* 31, 2041–2054. doi: 10.3987/COM-90-5563
- Miftakhov, M. S., Valeev, F. A., and Gaisina, I. N. (1994). Levoglucosenone: the properties, reactions, and use in fine organic synthesis. *Russ. Chem. Rev.* 63, 869. doi: 10.1070/RC1994v063n10ABEH000123
- Montgomery, D. (2008). *Design and Analysis of Experiments*. Hoboken, NJ: John Wiley & Sons.
- Novozymes (2015). <http://www.novozymes.com>
- Paris, C., Moliner, M., and Corma, A. (2013). Metal-containing zeolites as efficient catalysts for the transformation of highly valuable chiral biomass-derived products. *Green Chem.* 15, 2101–2109. doi: 10.1039/c3gc40267c
- Partridge, J., Harper, N., Moore, B. D., and Halling, P. J. (2001). “Control of acid-base conditions in low-water media,” in *Enzymes in Nonaqueous Solvents*,

- Vol. 15 of *Methods in Biotechnology*, eds E. Vulfson, P. Halling, and H. Holland (Totowa, NJ: Humana Press), 227–234.
- Perot, G., and Guisnet, M. (1990). Advantages and disadvantages of zeolites as catalysts in organic chemistry. *J. Mol. Catal.* 61, 173–196. doi: 10.1016/0304-5102(90)85154-A
- Sarotti, A. M., Zanardi, M. M., Spanevello, R. A., and Suarez, A. G. (2012). Recent applications of levoglucosenone as chiral synthon. *Curr. Org. Synth.* 9, 439–459. doi: 10.2174/157017912802651401
- Wei, X., Wang, Z., Wu, Y., Yu, Z., Jin, J., and Wu, K. (2014). Fast pyrolysis of cellulose with solid acid catalysts for levoglucosenone. *J. Anal. Appl. Pyrolysis* 107, 150–154. doi: 10.1016/j.jaap.2014.02.015

**Conflict of Interest Statement:** The authors declare that the research was conducted in the absence of any commercial or financial relationships that could be construed as a potential conflict of interest.

Copyright © 2016 Teixeira, Flourat, Peru, Brunissen and Allais. This is an open-access article distributed under the terms of the Creative Commons Attribution License (CC BY). The use, distribution or reproduction in other forums is permitted, provided the original author(s) or licensor are credited and that the original publication in this journal is cited, in accordance with accepted academic practice. No use, distribution or reproduction is permitted which does not comply with these terms.